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~~#3- ternary ionic compounds Naming~~

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~~Mr Pauller Type I Binary Ionic~~

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Binary and Ternary Ionic Compounds

Directions: Complete the following table.

Cation	Anion	Type	Formula	Unit Name
Mg ²⁺	Cl ¹⁻	Binary	MgCl ₂	Magnesium Chloride
NH ₄ ¹⁺	S ²⁻	Ternary	(NH ₄) ₂ S	Ammonium sulfide
Na ³⁺	SO ₃ ²⁻	Ternary	Na ₂ SO ₃	Sodium sulfite
Al ³⁺	N ³⁻	Binary	Al ₃ N	

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3 Aluminum Nitrate Ca²⁺ N³⁻ Binary
Ca₃ N₂ Calcium Nitride K₄ CrO₂⁻
Ternary K₂ CrO₄ Potassium
chromate Ba²⁺ CO₃²⁻ Ternary Ba₂
(CO₃)₂ ...

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File Type Binary Ionic Compounds –
Compounds with monatomic ions in it,
a metallic ion and a nonmetallic ion.

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This allows only two types of atoms in the formula. Ex: Rb_2O Ternary Ionic Compounds – Compounds with at least one polyatomic ion in it. This allows three or more types of atoms in the formula.

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A – side 2 1. Copper (II) sulfate 2.
Potassium chlorate 3. Ammonium
hydroxide 4. Potassium carbonate 5.
Sodium sulfate 6. Potassium acetate
7. Ammonium chloride 8. Magnesium
hydroxide 9. Zinc nitrate 10.
Magnesium hydroxide 11. Calcium
carbonate 12. Potassium hydroxide

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13. Calcium hydroxide 14.

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Compounds – Compounds with
monatomic ions in it, a metallic ion and
a nonmetallic ion. This allows only two
types of atoms in the formula. Ex:

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Rb₂O Ternary Ionic Compounds –
Compounds with at least one
polyatomic ion in it.

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File Type

Ternary Ionic CompoundsWWWoorr
kkkssshheeeeeett AAA — d ssiiddeee

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222. Ternary Ionic Compounds WWWo
oorrrkkkssshhheeeeett AAA —d
ssiiddeee 222. Directions: Circle or
underline the polyatomic ion in each
compound then write the correct name
or formula. Don't forget to use roman
numerals on transition metals. 1.

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kkkssshhheeeett AAA — d ...

Ternary Ionic Compounds Worksheet
Ternary Ionic Compounds Worksheet
A – side 2 1. Copper (II) sulfate 2.
Potassium chlorate 3. Ammonium
hydroxide 4. Potassium carbonate 5.
Sodium sulfate 6. Potassium acetate

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7. Ammonium chloride 8. Magnesium hydroxide 9. Zinc nitrate 10. Magnesium hydroxide 11. Calcium carbonate 12. Potassium hydroxide 13. Calcium hydroxide 14.

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- A – side 2
1. Copper (II) sulfate
 2. Potassium chlorate
 3. Ammonium hydroxide
 4. Potassium carbonate
 5. Sodium sulfate
 6. Potassium acetate
 7. Ammonium chloride
 8. Magnesium hydroxide
 9. Zinc nitrate
 10. Magnesium hydroxide
 11. Calcium

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carbonate 12. Potassium hydroxide
13. Calcium hydroxide 14. Potassium
carbonate 15. Aluminum phosphate
16.

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page 1 of the worksheet. This is due by Friday, October 30th. This information you will find in the Nomenclature slides and notes.

Ternary Ionic Compounds with
Polyatomic Ions 10/27/2020

Worksheet: More Binary and Name

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CHEMISTRY: A Study of Matter ©

2004, GPB 6.7a KEY Write the

formula of the compound formed from
the following pairs of ions: 1. Li^+ ,

O^{2-} _____ 2. Ag^+ , OH^- _____ 3. Mg^{2+} ,

SO_4^{2-} _____ 4. NH_4^+ , CO_3^{2-}

_____ 5. Ca^{2+} , P^{3-} _____ 6. Sn^{4+} ,

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